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Pulverulent reactive composition and process for the
purification of a gas

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The invention relates to the purification of
5 gases.

It relates more particularly to a reactive composition based on sodium bicarbonate which can be used for the purification of gases.

Human activities generate large amounts of
10 gases contaminated by toxic substances. Hydrogen chloride, hydrogen fluoride, sulphur oxides, nitrogen oxides, dioxins and furans are examples of toxic substances which are frequently found in these gases. Variable amounts of them are found in particular in the
15 flue gases generated by plants for the incineration of domestic or hospital waste and in the flue gases generated by the combustion of fossil fuels, in particular in thermal power stations for the generation of electricity and in centralized district heating
20 plants. These flue gases generally have to be freed from these toxic substances before being discharged to the atmosphere.

Sub C2
25 The Neutrec® process [Solvay (Société Anonyme)] is an efficient process for purifying gases. According to this known process, sodium bicarbonate, in the form of a powder, is injected into the gas and the gas thus treated is subsequently conveyed to a filter for removal of the dust therefrom (Solvay S.A., booklet Br. 1566a-B-1-0396).

30 Sodium bicarbonate powder has a natural tendency to cake, which constitutes a disadvantage. The addition of silica thereto has been contemplated in order to combat this disadvantageous property of sodium bicarbonate (Klein Kurt - "Grundlagen und Anwendungen
35 einer durch Flammenhydrolyse gewonnenen Kieselsäure: Teil 4: Aerosil zur Verbesserung des Fliessverhaltens pulverförmiger Substanzen" [Principles and applications of a silica produced by flame hydrolysis: Part 4: Aerosil for the improvement of the flow characteristics

of pulverulent substances] - Seifen-Ole-Fette-Wachse -
20 Nov. 1969, p. 849-858). However, sodium bicarbonate
to which silica has been added has not proved to be
very satisfactory in the purification of gases
5 comprising hydrogen chloride.

The invention overcomes this disadvantage by
providing a pulverulent reactive composition comprising
sodium bicarbonate which exhibits good resistance to
caking and satisfactory effectiveness in purifying a
10 gas.

The invention consequently relates to a solid
pulverulent reactive composition for the purification
of a gas, the said composition comprising sodium
15 bicarbonate and a caking inhibitor for sodium
bicarbonate and being characterized in that the
inhibitor comprises lignite coke and/or a magnesium
compound comprising magnesium (hydr)oxide.

Lignite coke is a product obtained by
carbonization of lignite, which is a solid fossil fuel
20 exhibiting a calorific value of less than 8,300 Btu/lb
(19.3 kJ/g) according to ASTM Standard D 388 (Ullmann's
Encyclopedia of Industrial Chemistry, 5th Edition,
Vol. A 7, 1986, pages 160-161).

The term "magnesium (hydr)oxide" is understood
25 to denote simultaneously magnesium oxide, magnesium
hydroxide or mixtures of magnesium oxide and magnesium
hydroxide. The magnesium compound advantageously
comprises basic magnesium carbonate of general formula
 $4\text{MgCO}_3 \cdot \text{Mg}(\text{OH})_2 \cdot 4\text{H}_2\text{O}$.

In addition to the sodium bicarbonate and the
inhibitor, the reactive composition according to the
invention can optionally comprise other constituents,
for example sodium monocarbonate or active charcoal.

The reactive composition according to the
35 invention preferably comprises more than 85%
(advantageously at least 90%) by weight of sodium
bicarbonate. Its content by weight of inhibitor is
preferably greater than 0.5% (advantageously at least
equal to 2%) of the weight of sodium bicarbonate. The

content by weight of inhibitor generally does not exceed 10% (preferably 7%) of the weight of the sodium bicarbonate. In the case where the inhibitor comprises lignite coke, the latter is preferably present in an 5 amount by weight of greater than 3% (advantageously at least equal to 5%) of the weight of the sodium bicarbonate. In the case where the inhibitor comprises a magnesium compound as defined above, the latter is preferably present in an amount by weight of greater 10 than 1% (advantageously at least equal to 2%) of the weight of the sodium bicarbonate.

In the case where the reactive composition according to the invention comprises sodium monocarbonate (of general formula Na_2CO_3), it is 15 desirable for its content by weight of sodium monocarbonate to be less than 2% (preferably at most equal to 1%) of the overall weight of sodium bicarbonate and sodium monocarbonate.

In an especially recommended embodiment of the 20 composition according to the invention, the latter exhibits a particle size defined by a mean particle diameter of less than 50 μm (preferably at most equal to 30 μm) and a particle size slope of less than 5 (preferably at most equal to 3). In this embodiment of 25 the invention, the mean diameter (D_m) and the particle size slope (σ) are defined by the following relationships:

$$D_m = \frac{\sum n_i \times D_i}{\sum n_i} \text{ [sic]}, \quad \sigma = \frac{D_{90} - D_{10}}{D_{50}} \text{ [sic]}$$

30 in which n_i denotes the frequency (by weight) of the particles of diameter D_i , and D_{90} (D_{50} and D_{10} respectively) represents the diameter at which 90% (50% and 10% respectively) of the particles of the reactive composition (expressed by weight) have a diameter of 35 less than D_{90} (D_{50} and D_{10} respectively). These particle

size parameters are defined by the method of analysis by laser radiation scattering using a Sympatec measuring device, Helos 12LA model, manufactured by Sympatec GmbH.

5 According to another recommended embodiment of the composition according to the invention, the latter is substantially devoid of silica. The phrase "substantially devoid of silica" is understood to mean that the amount of silica in the reactive composition 10 is insufficient to have a perceptible influence on the caking of the sodium bicarbonate, in the presence of atmospheric air, at a temperature of 20°C and at standard atmospheric pressure. The composition according to the invention is preferably entirely 15 devoid of silica. Everything else being equal, the composition in accordance with this embodiment of the invention exhibits optimum effectiveness as purification agent for gases.

20 The reactive composition according to the invention is applied as agent for the purification of gases contaminated by hydrogen chloride, hydrogen fluoride, sulphur oxides (mainly sulphur dioxide), nitrogen oxides (mainly nitric oxide NO and nitrogen peroxide NO₂), dioxins and furans. It is especially 25 advantageously applied in the purification of the flue gases generated by incinerators of municipal waste or hospital waste.

30 The invention also relates to a process for the purification of a gas, according to which a reactive composition comprising sodium bicarbonate is introduced into the gas and the gas is subsequently subjected to removal of dust, the process being characterized in that the reactive composition is substantially devoid of silica.

35 In the process according to the invention, the reactive composition is introduced in the solid state into the gas. The temperature of the gas is generally greater than 100°C (preferably greater than 125°C) during the introduction of the reactive composition. It

is recommended that the temperature of the gas should not exceed 800°C, preferably 600°C. Temperatures of 140 to 250°C are highly suitable. The reactive composition is generally introduced into a stream of gas moving in 5 a reaction chamber. The contaminants of the gas are, in the reaction chamber, adsorbed on the sodium bicarbonate particles (in the case of dioxins or furans) or react with the latter to form solid waste (for example, sodium chloride or fluoride, sodium 10 sulphate or sodium nitrite and nitrate, depending on whether the contaminants of the gas comprise hydrogen chloride, hydrogen fluoride, sulphur oxides or nitrogen oxides). The function of the removal of dust from the gas is to extract the solid waste thus formed 15 therefrom. Dust removal can be carried out by any appropriate known means, for example by mechanical separation in a cyclone, by filtration through a filter cloth or by electrostatic separation. Filtration through a filter cloth is preferred.

20 In accordance with the invention, it has been found that reactive compositions comprising sodium bicarbonate which are substantially devoid of silica are more effective in the purification of gases than sodium bicarbonate compositions comprising silica. This 25 improved effectiveness of the compositions according to the invention with respect to those comprising silica becomes evident mainly in the case where the removal of dust is carried out by means of a filter cloth. Although not wishing to be bound by a theoretical 30 explanation, the inventors believe that this greater effectiveness of the silica-free compositions can be attributed to the fact that these compositions adhere better to the filter cloth than the silica-comprising compositions.

35 In an advantageous embodiment of the process according to the invention, the reactive composition which is introduced into the gas is in accordance with the reactive composition according to the invention defined above and comprises, for this purpose, lignite

coke and/or a magnesium compound comprising magnesium (hydr)oxide.

The process according to the invention is especially advantageously applied in the purification 5 of a flue gas originating from the incineration of municipal waste or hospital waste, this waste generally comprising chlorinated compounds and metal chlorides capable of generating hydrogen chloride during incineration. This waste generally also comprises heavy 10 metals and sulphur-comprising waste, in particular sulphur dioxide, which are found at least partly in the flue gas. In this specific application of the process according to the invention, the solid product which is collected from the removal of dust consequently usually 15 comprises, in addition to sodium chloride, heavy metals in the metallic or combined state, as well as sodium carbonate and sodium sulphate. This solid product can be treated in the way set out in International Application WO 93/04983 [Solvay (Société Anonyme)].

20 The process according to the invention is also applied in the purification of flue gases generated by the combustion of fossil fuels (natural gas, liquid petroleum derivatives, coal), these flue gases being contaminated by sulphur dioxide and nitrogen oxides.

25 Furthermore, the process according to the invention is applied in the purification of fuel gases obtained by coal gasification, these gases generally being contaminated by hydrogen chloride, hydrogen fluoride and sulphur dioxide.

30 The advantage of the invention will emerge from the description of the following examples, with reference to the appended drawings.

Figure 1 diagrammatically shows a stack of bags comprising a reactive composition;

35 Figure 2 diagrammatically shows a device used to define the mobility of a pulverulent reactive composition.

In these figures, the same reference numbers denote identical components.

First series of tests.

Examples 1 to 6 relate to storage tests on reactive compositions in accordance with the invention, with the aim of assessing their resistance to caking.

5 To this end, in each of these examples, a solid and pulverulent reactive composition was bagged up in 15 polyethylene bags weighing 40 kg, which bags were hermetically sealed. The 15 bags were stacked on a support 7, in the way represented in Figure 1, so as to 10 form five rows (1, 2, 3, 4, 5) of three bags 6, and the stack of bags was stored in a warehouse with normal ventilation which is maintained at ambient temperature. After storage, the bags were opened, samples were 15 withdrawn therefrom in a random manner and two tests were carried out on the samples withdrawn. A first test served to define the tendency of the composition to cake. The second test served to evaluate the mobility 20 of the reactive composition, that is to say its ability to flow freely.

For the test targeted at defining the tendency to cake, the bags were poured out onto a graded screen with rectangular mesh openings of 12 x 19 mm and the degree of caking of the powder was defined by the relationship

25
$$D = \frac{\text{Amount by weight of agglomerates retained on the screen}}{\text{Total weight of powder poured onto the screen}} \times 100$$

Total weight of powder poured onto the screen

For the test targeted at defining the mobility of the reactive composition, use was made of the device represented diagrammatically in Figure 2. The device 30 comprises a sieve 9, exhibiting a mesh size of 710 μm , positioned above a vertical cylinder 10 with a diameter of 50 mm. For the test, the powder was poured through the sieve, the powder was collected on the top horizontal face 11 of the cylinder 10 and the maximum 35 height of the cone of powder 12 formed on the face 11 of the cylinder 10 was measured. According to this test, the mobility of the powder increases as the height of the cone 12 decreases.

Example 1

In this example, use was made of a reactive composition comprising milled and screened sodium bicarbonate, 0.48% by weight of silica and 4.6% by weight of lignite coke (the contents of silica and of lignite coke are expressed with respect to the weight of sodium bicarbonate). The screening of the sodium bicarbonate was adjusted so that the latter is in the form of particles not exceeding 13 μm in diameter, the reactive composition exhibiting a particle size defined by the following characteristics (defined above), expressed in μm :

D₁₀ [sic] = 7.0

D₅₀ [sic] = 29.7

D₉₀ [sic] = 70.3

After storage for three months, the composition was subjected to the two tests defined above. The following results were obtained:

Tendency to cake (test on three samples) :

Sample No. 1: 0.50%

Sample No. 2: 2.98%

Sample No. 3: 0.11%

Mobility (test on five samples) : Sample No. 1: 40 mm

Sample No. 2: 36 mm

Sample No. 3: 40 mm

Sample No. 4: 39 mm

Sample No. 5: 38 mm

Mean : 39 mm

Example 2

The tests of Example 1 were repeated with a reactive composition comprising milled and screened sodium bicarbonate, 1.89% by weight of basic magnesium carbonate and 5% by weight of lignite coke (the contents of basic magnesium carbonate and of lignite coke are expressed with respect to the weight of sodium bicarbonate [lacuna]). The screening of the sodium bicarbonate was adjusted as in Example 1, so that it is in the form of particles not exceeding 13 μm in diameter, the reactive composition exhibiting a

particle size defined by the following characteristics (defined above), expressed in μm :

5	D10 [sic] = 6.6
	D50 [sic] = 33.7
	D90 [sic] = 75.4

After storage for three months, the following results were obtained:

Tendency to cake (test on three samples): 0%

10 Mobility (test on five samples): Sample No. 1: 34 mm
Sample No. 2: 38 mm
Sample No. 3: 37 mm
Sample No. 4: 36 mm
Sample No. 5: 39 mm

Example 3

The tests of Example 1 were repeated with a reactive composition comprising milled and screened sodium bicarbonate and 5.1% by weight of lignite coke, the content of lignite coke being expressed with respect to the weight of sodium bicarbonate. The screening of the sodium bicarbonate was adjusted as in Example 1, so that it is in the form of particles not exceeding 13 μm in diameter, the reactive composition exhibiting a particle size defined by the following characteristics (defined above), expressed in μm :

D10 [sic] = 7.0
D50 [sic] = 35.1
D90 [sic] = 85.0

30 After storage for three months, the following results were obtained:

Tendency to cake (test on three samples): 0%

Mobility (test on five samples): Sample No. 1: 37 mm
 Sample No. 2: 38 mm
 Sample No. 3: 41 mm
 Sample No. 4: 40 mm
 Sample No. 5: 38 mm
 Mean : 39 mm

The preceding examples show that the reactive

compositions in accordance with the invention correctly endure storage for several months. A comparison of the results of Examples 2 and 3 with those of Example 1 furthermore show [sic] that the absence of silica in 5 the reactive composition is not harmful to its ability to be stored.

Examples 4 to 6

In Examples 4 to 6, the tests of Example 1 to 3 respectively were repeated with a storage time of six 10 months. The characteristics of the compositions are given in Table 1 below.

Table 1

	Examples [sic] No.		
	4	5	6
Silica (%)	0.5		
Basic magnesium carbonate (%)		2	
Lignite coke (%)	5	5	5
D10 [sic] (μm)	7.6	12.3	7.7
D50 [sic] (μm)	30.0	41.2	36.7
D90 [sic] (μm)	69.1	83.4	79.4

15 The results obtained after storage for six months are given in Table 2 below.

Table 2

	Examples [sic] No.		
	4	5	6
Tendency to cake			
Sample No. 1	0	0	0
Sample No. 2	3.2	0	0
Sample No. 3	3.1	0	0
Sample No. 4	1.8	0	0
Sample No. 5	0	0	0
Mobility			
Sample No. 1	43	29	43
Sample No. 2	41	30	38
Sample No. 3	46	29	43.5
Sample No. 4	44	28	45
Sample No. 5	43	30	41

Examples 4 to 6 confirm the results of Examples 1 to 3 by demonstrating the excellent ability of the silica-free reactive compositions according to the invention.

5 Second series of tests

Examples 7 to 10 relate to tests carried out with the aim of measuring the effectiveness of reactive compositions in purifying a gas from hydrogen chloride.

10 The gas treated in each test was a flue gas originating from an incinerator of domestic waste comprising hydrogen chloride and sulphur dioxide. An at least sufficient amount of a reactive composition comprising sodium bicarbonate was introduced into the flue gas to bring its residual content of hydrogen

15 chloride below 50 mg/Sm³ (European Standard 89/369/EEC) or below 10 mg/Sm³ (European Standard 94/67/EEC or German Standard 17.BIm SchV). After addition of the reactive composition, the flue gas was filtered through a filter cloth to remove dust therefrom.

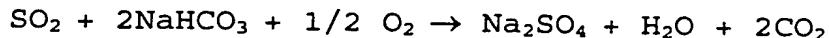
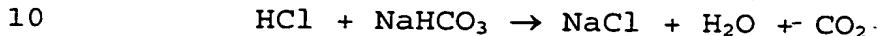
20 Example 7 (in accordance with the invention)

In this example, the reactive composition employed consisted essentially of sodium bicarbonate, without additive. In particular, the reactive composition was devoid of silica.

25 The test lasted 390 minutes. During the test, the flow rate of the flue gas, the throughput of the reactive composition introduced into the flue gas and the contents of hydrogen chloride and of sulphur dioxide in the flue gas were continuously measured, respectively upstream of the addition of the reactive composition and downstream of the filter cloth. From these measurements, the stoichiometric ratio (S.R.) of the amount of sodium bicarbonate actually employed to the stoichiometric amount required, on the one hand, 30 and the degree of purification from hydrogen chloride, the latter being defined by the relationship

$$\tau = \frac{HCl_i - HCl_f}{HCl_i} \times 100,$$

where HCl_i denotes the content of hydrogen chloride in the flue gas upstream of the addition of the reactive composition and HCl_f denotes the content of hydrogen chloride in the flue gas downstream of the said addition [sic], on the other hand, were calculated. In the test, the stoichiometric amount of sodium bicarbonate is that required to remove the hydrogen chloride and the sulphur dioxide from the flue gas, according to the following theoretical reactions:



The results of the test (arithmetic mean over the 390 minutes) are recorded below:

Flue gas

15	Flow rate (Sm^3/h)	:	2378
	HCl_i (mg/Sm^3)	:	1530
	HCl_f (mg/Sm^3)	:	9

Reactive composition:

20	$NaHCO_3$ throughput (kg/h)	:	13
	S.R.	:	1.49
	Degree of purification (%)	:	99.4

Example 8 (not in accordance with the invention)

The test of Example 7 was repeated with a reactive composition composed of sodium bicarbonate and 25 silica (0.5 g of silica per 100 g of sodium bicarbonate). The results of the test (which lasted 360 minutes) are given below.

Flue gas

30	Flow rate (Sm^3/h)	:	1697
	HCl_i (mg/Sm^3)	:	2018
	HCl_f (mg/Sm^3)	:	39

Reactive composition:

35	$NaHCO_3$ throughput (kg/h)	:	26
	S.R.	:	3.07
	Degree of purification (%)	:	98.1

A comparison of the results of Example 7 (in accordance with the invention) with those of Example 8 (not in accordance with the invention) immediately reveals the advantage of avoiding, in accordance with

the invention, the presence of silica in the reactive composition.

Example 9 (in accordance with the invention)

The test of Example 7 was repeated with a reactive composition in accordance with the invention which is devoid of silica and is composed of a homogeneous mixture of sodium bicarbonate and basic magnesium carbonate (2 g per 100 g of sodium bicarbonate). The results of the test (which lasted 67 hours) are given below.

Flue gas

Flow rate (Sm ³ /h)	:	24,000
HCl _i (mg/Sm ³)	:	1060
HCl _f (mg/Sm ³)	:	32

Reactive composition:

NaHCO ₃ throughput (kg/h)	:	63.7
S.R.	:	1.11
Degree of purification (%)	:	99.0

Example 10

The test of Example 7 was repeated with a reactive composition in accordance with the invention which is devoid of silica and is composed of a homogeneous mixture of sodium bicarbonate and lignite coke (5 g per 100 g of sodium bicarbonate). The results of the test (which lasted 81 hours) are given below.

Flue gas

Flow rate (Sm ³ /h)	:	24,000
HCl _i (mg/Sm ³)	:	925
HCl _f (mg/Sm ³)	:	46

Reactive composition:

NaHCO ₃ throughput (kg/h)	:	63.8
S.R.	:	1.09
Degree of purification (%)	:	>99.9

Examples 9 and 10 show the positive influence of the basic magnesium carbonate and lignite coke on the effectiveness of the reactive composition.